REMARKS

Applicants amended Tables 4 and 5 of the specification as filed to correct typographical errors in the units for density and molecular volume. As disclosed in paragraphs [00121] and [0124] of the specification as filed, the density of the elements and element oxides in Tables 4 and 5 can be found on specific pages of Section 4 of the CRC Handbook of Chemistry and Physics, 81^{rst} Edition, D.R. Lide, Ed., CRC Press, Inc., 2000-2001, which identifies the unit for density of these elements and element oxides listed in these tables as 'g cm^{-3}' (see for example in the APPENDIX the attached Page 4-39 of such publication on Page 6 of this paper). By this amendment, Applicants corrected the unit for density in Tables 4 and 5 by replacing 'cm^3/g' with its proper unit 'g/cm^3'. Furthermore, Applicants corrected the spelling of the unit for molecular volume in Table 5 from 'Cm^3/mol' to 'cm^3/mol'.

Applicants submit that no new matter was introduced to the specification by way of this amendment to Tables 4 and 5; and that the amendment to the specification does not impact the scope of the claims and thus does not raise new issues that would require further consideration and/or search. In addition, as the editorial changes to the specification are minimal, this amendment would not involve materially added work on the part of the Office.

This amendment under 37 C.F.R. 1.312 is submitted on or before the payment of the issue fee, and Applicants respectfully request its entry for the reasons stated above.

Should a petition for extension of time be necessary in order for this paper to be deemed timely filed, please consider this a petition therefor. If any fee is due, please appropriately charge such fee to Deposit Account Number 16-1575 of ConocoPhillips Company, Houston, Texas.

Should there be any remaining issue, the Examiner is invited to call the undersigned at (281) 293-4751.

Respectfully submitted,

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CRC Handbook of Chemistry and Physics

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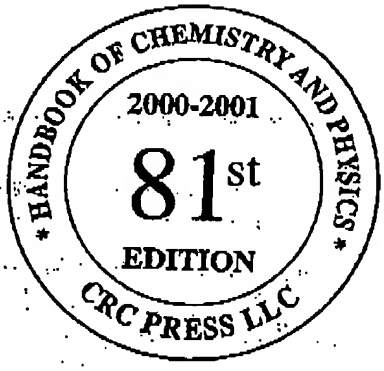

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PHYSICAL CONSTANTS OF INORGANIC COMPOUNDS

No.	Name, Formula	CAS RN, Mol. Wt.	Physical Form	mp/°C, dm/g cm^{-3}	bp/°C, Other Data	Solubility
1	Actinium, Ac	7440-34-8, 227	silv metal; cub	1051, 10	3198, a	
2	Actinium bromide, $AcBr_3$	33689-81-5, 487	wh hex cry	5.85		s H_2O
3	Actinium chloride, $AcCl_3$	22986-54-5, 333	wh hex cry	4.81		
4	Actinium fluoride, AcF_3	33689-80-4, 284	wh hex cry	7.88		i H_2O
5	Actinium iodide, AcI_3	33689-82-6, 608	wh cry			s H_2O
6	Actinium oxide, Ac_2O_3	12002-61-8, 502	wh hex cry	1977, 9.19		i H_2O
7	Aluminum, Al	7429-90-5, 26.982	silv-wh metal; cub cry	660.32, 2.70	2519, a,c,d,e	i H_2O; s acid, alk
8	Aluminum ammonium sulfate, $AlNH_4(SO_4)_2$	7784-25-0, 237.148	wh powder			sl H_2O; i EtOH
9	Aluminum ammonium sulfate dodecahydrate, $AlNH_4(SO_4)_2{\cdot}12H_2O$	7784-26-1, 453.331	col cry or powder	94.5, 1.65	dec >280	s H_2O; i EtOH
10	Aluminum antimonide, AlSb	25152-52-7, 148.742	cub cry	1065, 4.26		
11	Aluminum arsenide, AlAs	22831-42-1, 101.903	oran cub cry; hyg	1740, 3.76		
12	Aluminum borate, $2Al_2O_3{\cdot}B_2O_3$	11121-16-7, 273.543	needles	≈1050		i H_2O
13	Aluminum boride, AlB_2	12041-50-8, 48.604	powder	dec >920, 3.19		s dil HCl
14	Aluminum borohydride, $Al(BH_4)_3$	16962-07-5, 71.510	flam liq	-64.5	44.5, a,d	reac H_2O
15	Aluminum bromate nonahydrate, $Al(BrO_3)_3{\cdot}9H_2O$	11126-81-1*, 572.828	wh hyg cry	62	dec >100	s H_2O
16	Aluminum bromide, $AlBr_3$	7727-15-3, 266.694	wh-yel monocl cry; hyg	97.5, 3.2	256, a,b,c,d,e	reac H_2O; s bz, tol
17	Aluminum bromide hexahydrate, $AlBr_3{\cdot}6H_2O$	7784-11-4, 374.785	col-yel hyg cry	90, 2.54		s H_2O, EtOH, CS_2
18	Aluminum carbide, Al_4C_3	1299-86-1, 143.959	yel hex cry	2100, 2.36	dec >2200	reac H_2O
19	Aluminum chlorate nonahydrate, $Al(ClO_3)_3{\cdot}9H_2O$	15477-33-5, 439.472	hyg cry			vs H_2O; s EtOH
20	Aluminum chloride, $AlCl_3$	7446-70-0, 133.340	wh hex cry or powder; s hyg	192.6, 2.48	a,b,c,e	reac H_2O; s bz, clo, chl
21	Aluminum chloride hexahydrate, $AlCl_3{\cdot}6H_2O$	7784-13-6, 241.431	col hyg cry	dec 100, 2.398		vs H_2O; s EtOH, eth
22	Aluminum diacetate, $Al(OH)(C_2H_3O_2)_2$	142-03-0, 162.079	wh amorp powder			i H_2O
23	Aluminum ethoxide, $Al(C_2H_5O)_3$	555-75-9, 162.165	liq, condenses to wh solid	140		reac H_2O; sl xyl
24	Aluminum fluoride, AlF_3	7784-18-1, 83.977	wh hex cry	≈2250 tp (220 MPa), 3.10	1276 sp, a,c,e	sl H_2O
25	Aluminum fluoride monohydrate, $AlF_3{\cdot}H_2O$	32287-65-3, 101.992	orth cry	2.17		sl H_2O
26	Aluminum fluoride trihydrate, $AlF_3{\cdot}3H_2O$	15098-87-0, 138.023	wh hyg cry	1.914		sl H_2O
27	Aluminum hexafluorosilicate nonahydrate, $Al_2(SiF_6)_3{\cdot}9H_2O$	17099-70-6, 642.329	hex prisms	dec >500		s H_2O
28	Aluminum hydride, AlH_3	7784-21-6, 30.006	col hex cry	dec >150	g	reac H_2O
29	Aluminum hydroxide, $Al(OH)_3$	21645-51-2, 78.004	wh amorp powder	2.42	f	i H_2O; s alk, acid
30	Aluminum hydroxychloride, $Al_2(OH)_5Cl{\cdot}2H_2O$	1327-41-9, 210.483	gl solid			s H_2O
31	Aluminum hypophosphite, $Al(H_2PO_2)_3$	7784-22-7, 221.948	cry powder	dec 220		i H_2O; s alk, acid
32	Aluminum iodide, AlI_3	7784-23-8, 407.695	wh leaflets	188.32, 3.98	382, a,b,c,d,e	reac H_2O
33	Aluminum iodide hexahydrate, $AlI_3{\cdot}6H_2O$	10090-53-6, 515.786	yel hyg cry powder			vs H_2O; s EtOH, eth
34	Aluminum lactate, $Al(C_3H_5O_3)_3$	18917-91-4, 294.195	powder			vs H_2O
35	Aluminum nitrate nonahydrate, $Al(NO_3)_3{\cdot}9H_2O$	7784-27-2, 375.134	wh hyg cry; monocl	73, 1.72	dec 135	vs H_2O, EtOH; i pyr
36	Aluminum nitride, AlN	24304-00-5, 40.989	blue-wh hex cry	3000, 3.255	a	reac H_2O
37	Aluminum oleate, $Al(C_{18}H_{33}O_2)_3$	688-37-9, 871.358	yel solid			i H_2O; s EtOH, bz
38	Aluminum phosphate, $AlPO_4$	7784-30-7, 121.953	wh rhomb plates	>1460, 2.56	a	i H_2O; sl acid